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Unit Operations of Chemical Engineering

SEVENTH EDITION

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2.26 (a) $70.5 \text{ lb}_m / \text{ft}^3$; $8.27 \times 10^{-7} \text{ in}^2 / \text{lb}_f$

(b) $\rho = (70.5 \text{ lb}_m / \text{ft}^3) \exp \left[\frac{8.27 \times 10^{-7} \text{ in}^2}{\text{lb}_f} \left| \frac{9 \times 10^6 \text{ N}}{\text{m}^2} \right| \frac{14.696 \text{ lb}_f / \text{in}^2}{1.01325 \times 10^5 \text{ N/m}^2} \right]$
 $= \frac{70.57 \text{ lb}_m}{\text{ft}^3} \left| \frac{35.3145 \text{ ft}^3}{\text{m}^3} \right| \frac{1 \text{ m}^3}{10^6 \text{ cm}^3} \left| \frac{1000 \text{ g}}{2.20462 \text{ lb}_m} \right| = 1.13 \text{ g/cm}^3$

(c) $\rho \left(\frac{\text{lb}_m}{\text{ft}^3} \right) = \rho' \frac{\text{g}}{\text{cm}^3} \left| \frac{1 \text{ lb}_m}{453.593 \text{ g}} \right| \frac{28.317 \text{ cm}^3}{1 \text{ ft}^3} = 62.43 \rho'$
 $\rho \left(\frac{\text{lb}_f}{\text{in}^2} \right) = \rho'' \frac{\text{N}}{\text{m}^2} \left| \frac{0.2248 \text{ lb}_f}{1 \text{ N}} \right| \frac{1^2 \text{ m}^2}{39.37^2 \text{ in}^2} = 1.45 \times 10^{-4} \rho''$
 $\Rightarrow 62.43 \rho' = 70.5 \exp \left[(8.27 \times 10^{-7}) (1.45 \times 10^{-4} \rho'') \right] \Rightarrow \rho' = 1.13 \exp (1.20 \times 10^{-10} \rho'')$
 $\rho'' = 9.00 \times 10^5 \text{ N/m}^2 \Rightarrow \rho' = 1.13 \exp (1.20 \times 10^{-10} \times 9.00 \times 10^5) = 1.13 \text{ g/cm}^3$

2.27 (a) $V(\text{cm}^3) = V'(\text{in}^3) \left| \frac{28.317 \text{ cm}^3}{1728 \text{ in}^3} \right| = 16.39V'$; $t(\text{s}) = 3600t'$ (hr)
 $\Rightarrow 16.39V'^n = \exp(3600t') \Rightarrow V' = 0.06102 \exp(3600t')$

(b) The t in the exponent has a coefficient of s^{-1} .

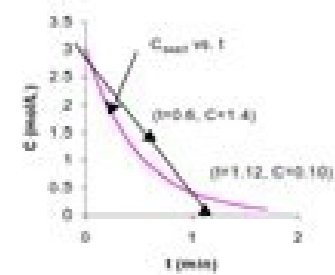
2.28 (a) 3.00 mol/L , 2.00 min^{-1}

(b) $t = 0 \Rightarrow C = 3.00 \exp(-2.00 \times 0) = 3.00 \text{ mol/L}$
 $t = 1 \Rightarrow C = 3.00 \exp(-2.00 \times 1) = 0.406 \text{ mol/L}$

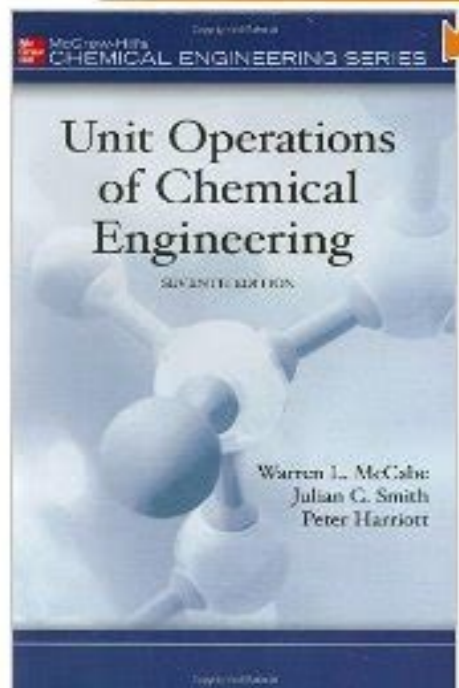
For $t = 0.6 \text{ min}$: $C_{\text{in}} = \frac{0.406 - 3.00}{1 - 0} (0.6 - 0) + 3.00 = 1.4 \text{ mol/L}$
 $C_{\text{out}} = 3.00 \exp(-2.00 \times 0.6) = 0.9 \text{ mol/L}$

For $C = 0.10 \text{ mol/L}$: $t_{\text{in}} = \frac{1 - 0}{0.406 - 3} (0.10 - 3.00) + 0 = 1.12 \text{ min}$
 $t_{\text{out}} = -\frac{1}{2.00} \ln \frac{C}{3.00} = -\frac{1}{2} \ln \frac{0.10}{3.00} = 1.70 \text{ min}$

(c)



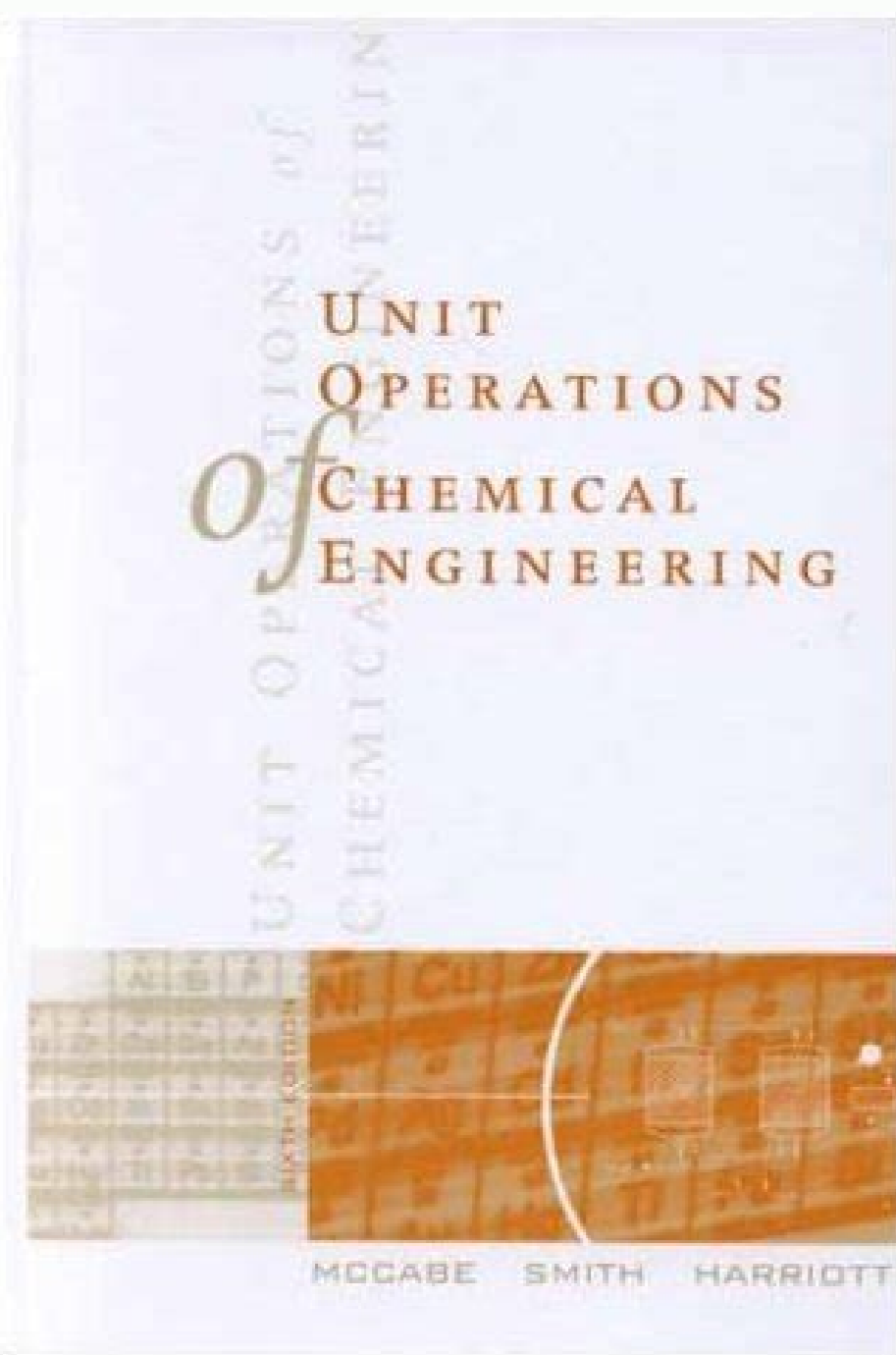
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